# Electrochemistry of Nitrogen and Boron Bi-element Incorporated Diamond Films

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**Abstract** (188 words)

Boron doped diamond (BDD) has been widely used in various electrochemical fields, due to its unique physical and chemical properties. However, the investigation of the electrochemistry of bi-element incorporated diamond, especially the variation of surface components after nitrogen incorporation into BDD and the corresponding electron transfer of inner-sphere and outer-sphere redox probes is still lacking. Here, the electrochemistry of nitrogen and boron bi-element incorporated diamond (NBD) is thus investigated in both inner and outer redox systems, namely in $[Fe(CN)_6]^{3-/4-}$ and $[Ru(NH_3)_6]^{3+/2+}$ solutions. On the NBD electrode, enhanced electrochemical responses are achieved for both outer- and inner-sphere redox reactions. Such enhancement originates from the enrichment of C=O groups and the increased amount of $sp^2$ species in the NBD film. Moreover, the nitrogen and boron atoms incorporated in diamond modulate the surface polarities and the electronic states of diamond. Based on the enhanced and stable capacitance of the NBD electrode, its electrochemical energy applications are explored by assembling its supercapacitor as a case study. This work reveals the influence of $sp^2$ species and oxygen-contained function group on the electrochemistry of bi-element incorporated diamond films and reveals their potential electrochemical applications.

**Keywords:** Electrochemistry, nitrogen and boron-doped diamond, $sp^2$-carbon

## 1. Introduction

The electrochemistry of doped diamond, especially p-type diamond or boron doped diamond (BDD) has been widely investigated in the past decades. As an excellent electrode material, BDD has been utilized for electrochemical sensors [1, 2], supercapacitor construction [3-5], $CO_2$ reduction [6, 7], nitrogen redox reaction (NRR) [8], and wastewater treatment [9, 10]. These applications originate from the unique features of a BDD electrode, such as its low background current, wide potential windows in different media, and long-term durability. Boron atoms doped in diamond realize the transformation of diamond from an intrinsic insulator to a semiconductor [11], and finally to a metal-like conductor once the boron doping level increases above to $10^{20}$ $cm^{-3}$ [12]. It is well-known that the diamond crystal structure, surface termination, and $sp^2$ species or $sp^2/sp^3$ ratio on the diamond surface play significant roles to determine the electrochemical features of a BDD electrode. For example, a rough BDD surface promotes the transformation of the Faradaic reactions from kinetic- to diffusion-control together with enhanced charge transfer rates [13, 14]. The surface (e.g., hydrogen, oxygen) terminations of a BDD film influence the kinetics of Faradaic reactions on the diamond surface because these terminations possess significant difference in their electronic structures [15-17]. In this regard, the electrochemistry of BDD films containing various amounts of $sp^2$ carbon has been also extensively studied, although their quality is much reduced and their background currents are much enhance [18]. To further boost the performance of diamond films in the fields of energy and catalysis applications, diamond composite structures have been designed, for instance to assemble battery-like supercapacitors by use of aligned carbon nanofiber coated BDD [4], to achieve an efficient methanol oxidation reaction (MOR) using nanoporous platinum particles coated BDD [19].

On the other hand, electrochemistry of n-type diamond, namely diamond films doped with nitrogen or phosphorus atoms has also attracted much attention. Nitrogen doping or incorporation into carbon materials has also been confirmed as a fruitful strategy to promote the electrocatalytic activity of these carbon materials. The pyridinic N atoms create Lewis basic sites that are actually regarded as the catalytic active sites [20, 21]. For example, a nitrogen doped diamond (NDD) film is proved to contain N-sp$^3$ components, namely electrocatalytic active sites [22, 23]. In this context, a NDD film exhibits high overpotential for the hydrogen evolution reaction (HER) and has been applied for highly efficient $CO_2$ reduction [24, 25].

We are interested in the electrochemistry of bi-element incorporated diamond films. Compared with diamond films doped with a single dopant, diamond films with dual dopants are expected to regulate the electronic structure of diamond materials and eventually exhibit faster electron transfer rates and more active sites for catalysis [26-28]. Such enhanced electrochemical performance stems from the synergistic effects of two different and incorporated atoms in the diamond film. One recent example is the application of nitrogen and boron co-doped diamond (NBD) film for the efficient $CO_2$ reduction [29]. The NBD film with optimized contents of nitrogen and boron dopants exhibited comparable performance toward oxygen reduction reaction (ORR) to the Pt/C catalyst, including a high current density for ORR and long-term durability of the system [30]. In spite of these successful catalytic applications of these bi-element incorporated diamond films, the electrochemistry of the NBD films has been seldom investigated. For example, the variation of surface components in the NBD films and their influence on the electron transfer rates of both inner-sphere and outer-sphere redox systems have not been clarified up to now. Moreover, reports about the applications of bi-element incorporated diamond films for energy storage are still missing in the literature, although BDD and its composites are shown to be promising electrode candidates for the assembly of supercapacitors [4, 31, 32]. Therefore, this contribution deals with the electrochemistry of the NBD films that are grown by a microwave plasma enhanced chemical vapor deposition (MPCVD) method. After the characterization of this NBD film with different techniques, its electrochemical responses are studied in both $[Fe(CN)_6]^{3-/4-}$ and $[Ru(NH_3)_6]^{3+/2+}$ redox systems, which are further compared with the BDD electrode. As a case study of the energy applications of these NBD films, a supercapacitor is assembled and investigated.

## 2. Experiment section

### *2.1 Materials synthesis and characterization*

The NBD and BDD films were grown on the Si (100) wafers using a MPCVD method [33-35]. The detailed growth parameters are listed in **Table S1**.

The SEM images of the as-grown NBD and BDD films were recorded with a field emission scanning electron microscope (FESEM, Zeiss ultra55, Germany). The transmission electron microscopy (TEM, FEI $G^2$ F20) was employed to characterize the defects the crystalline defects in the as-grown NBD and BDD films. The surface chemical composition of these as-grown diamond films was analyzed by X-ray photoelectron spectroscopy (S-probe ESCA SSX-100s, Surface Science Instruments, USA) with an Al Kα radiation of 200 W. The survey spectra were measured from 0 to 1200 eV with a resolution of 1 eV at a spot size of 800 $\mu m^2$. The high resolution spectra were collected with a resolution of 0.1 eV at a spot size of 300 $\mu m^2$. The Raman spectra of the as-grown diamond films were collected on a homemade Raman Instrument equipped with a 532-nm laser. A time-of-flight secondary ion mass spectrometer (ToF-SIMS IV, ION-TOF GmbH, Germany) was used to map the dopants in these as-grown diamond films, such as the contents of nitrogen and boron atoms in the NBD films as well as boron atom in the BDD film. For these mapping experiments, a 25-keV $Bi^+$ primary ion beam was employed to bombard the diamond surface within an area of 300 × 300 $\mu m^2$.

## *2.2 Electrochemical measurements*

Electrochemical measurements of the as-grown NBD and BDD films were conducted on a CHI660e workstation (Shanghai Chenhua Inc., China) using a three-electrode cell, where an Ag/AgCl (3M KCl) electrode acted as reference electrode, a Pt wire as counter electrode, a NBD film or a BDD film as working electrode. The geometric area of a working electrode was 0.05 $cm^2$. To investigate the electrochemical performance of the NBD and BDD films, their cyclic voltammograms (CVs) were recorded in either 1 mM $K_3[Fe(CN)_6]$ or $[Ru(NH_3)_6]Cl_3$ dissolved in 1 M KCl aqueous solution. The investigation of the pseudocapacitive behavior of the post-treated NBD and BDD films was carried out by means of cyclic voltammetry at different scan rates and by means of the galvanostatic charge/discharge (GCD) method at different current densities. The post-treatment was conducted in the mixture of $H_2SO_4$ and $HNO_3$ (V/V = 3:1) for 30 min. In this way, these diamond films were found to exhibit better wettability in the electrolytes. The electrolyte used for the assemble of a supercapacitor was 0.05 M $K_3Fe(CN)_6/K_4Fe(CN)_6$ dissolved in 1.0 M $Na_2SO_4$ solution. The specific capacitances were calculated according to the reported methods [4, 36]. The calculation of the contribution of the capacitive current was based on the equation of $i\,(V) = k_1 v + k_2 v^{1/2}$ [37, 38]. Here, $i\,(V)$ is the related current at the potential of V, $v$ is scan rate, $k_2 v^{1/2}$ and $k_1 v$ are related to diffusion-controlled and capacitive-controlled, respectively. Note that the capacitive current can be also evaluated directly from the cyclic voltammograms (CVs) or the GCD curves in the blank solutions (namely those containing only supporting electrolytes).

## 3. Results and discussion

### *3.1 Characterization of the NBD films*

The morphologies of the as-grown BDD and NBD films were analyzed by electron microscopy. From the typical SEM images of the NBD (**Figure 1a**) and BDD (**Figure S1**) films, one can see clearly that these films exhibit typical and similar morphology to that of polycrystalline diamond films. Their grain sizes are in the range of 0.4 - 1.2 μm. The cross-sectional SEM images of the NBD and BDD films (**Figure S2**) reveal their thickness to be about 1.5 μm. To check out crystalline defects on these films, the TEM images of the NBD film were recorded (**Figure 1b**, **1c**), where twin boundaries and stacking faults are observed. The presence of these defects is caused by the incorporation of both nitrogen and boron atoms into the diamond film. At selected locations for TEM imaging experiments, it seems to be that the crystalline defects of the NBD film are reduced, compared to the BDD film (**Figure S1**). Meanwhile, the crystalline quality of a NBD film seems to be improved and the {100} texture of diamond is promoted [39]. In a high-resolution TEM (HRTEM) image of a NBD film (**Figure 1c**), the atomic structure of the NBD film can be clearly seen along the [01-1] zone axis. According to the inset of fast Fourier transformation (FFT), the diffraction spots reveal spacings of 0.206 and 0.18 nm. These spacings correspond to the (111) and (200) planes of diamond phase, respectively. Consequently, the as-grown NBD and BDD films exhibit high crystallinity.

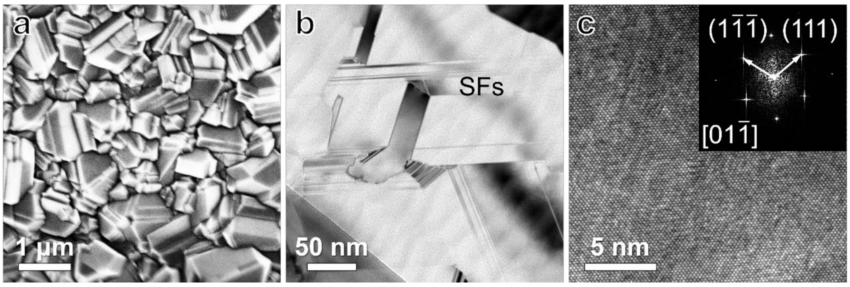


**Figure 1**. (a) SEM, (b) low-magnification TEM and (c) HRTEM images of a NBD film. The inset in (c) is the corresponding fast Fourier transformation (FFT) of the HRTEM image.

The Raman spectra of the as-grown NBD and BDD films were also recorded (**Figure 2a**). In both spectra, the typical Raman peak of diamond is seen around 1320 $cm^{-1}$. The Lorentzian peak located around 1200 $cm^{-1}$ is classified as the symmetry breaking of the diamond lattice. The spectrum of a NBD film displays a finite blueshift, resulting from a higher bond energy when nitrogen and boron atoms are bi-element incorporated in the diamond lattice [40]. Moreover, both Raman spectra of the as-grown NBD and BDD films reveal a broad peak around at 1580 $cm^{-1}$. It is known as the G band that results from the bond stretching of $sp^2$ atoms in both rings and chains.

To determine the surface chemical bonding of the as-grown NBD and BDD films, their survey and C1s XPS spectra were recorded and compared (**Figure S3 and Figure 2b**). In both XPS spectra, four peaks centered at 282.8, 284.1, 285.2, and 286.3 eV are attributed to $sp^2$ and $sp^3$ hybridized carbon species, as well as carbon bonded to oxygen as C–O and as C=O, respectively [41, 42]. The peak located around 284.8 eV in the

NBD film corresponds to carbon bonded to nitrogen. Furthermore, it can be seen that the NBD film reveals a substantially higher $sp^2/sp^3$ ratio than the BDD film together with an increased fraction of carbon in C=O and C-O bonds (**Table S2**). The nitrogen atoms incorporated into diamond are expected to be three-fold coordinated in the amorphous/disordered regions with the remaining electrons in a lone pair configuration. In other words, the nitrogen atoms incorporated into diamond promotes the formation of sp2 carbon [43, 44]. Under such conditions, nitrogen incorporation into diamond tends to change the bonding, instead of being assimilated by the diamond lattice that is not the intrinsic of electronic dopant. **Figure 2c** shows that the B1s peak of the NBD film is shifted to a lower binding energy compared to the B1s peak of the BDD film. Presumably, this is due to the formation of B–N bonds in the NBD film. In the N1s XPS spectrum of the NBD film (**Figure 2d**), two peaks are detected at 399.7 and 397.5 eV, which are attributed to nitrogen atoms bonded to carbon and to boron, respectively [25, 29, 30]. In the XPS spectrum of the BDD film, no N1s peak was detected. The ratios of nitrogen to carbon and boron to carbon were estimated from the high resolution XPS spectra of the NBD film. They are 0.013 and 0.008, respectively. Similarly, the ratio of boron to carbon in the BDD film is 0.008. Furthermore, boron atoms are found to be homogeneously and uniformly distributed throughout the film, as confirmed from secondary ion mass spectrometry (SIMS) mappings of doped boron atoms in the NBD and BDD films (**Figure 2e, 2f, S4**). Surprisingly, nitrogen was not detectable with the current SIMS setup, due to isobaric interferences from carbon species with similar mass-to-charge ratios (*m/z*) (**Figure S4 a-b**). Meanwhile, the content of incorporated nitrogen in the NBD film under investigation is presumably not very high (e.g., less than $10^{18}$ atoms $cm^{-3}$) and close to or at the detection limit of our SIMS setup. In the future, a better primary beam intensity and improved vacuum will help to achieve improved detection limits for incorporated nitrogen atoms in these NBD films.

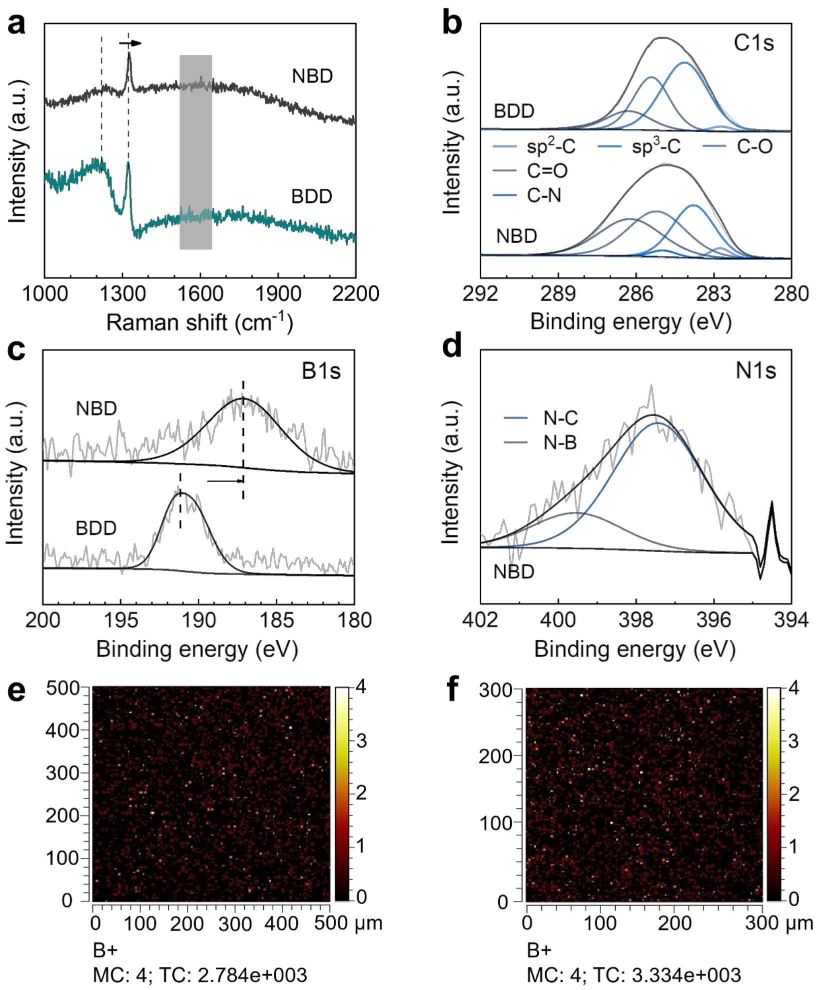

**Figure 2**. (a) Raman spectra of the NBD and BDD films; their C1s (b), B1s (c) and N1s (d) XPS spectra; SIMS mapping of boron atoms doped in the NBD film directly (e) and after 30 sec sputtering with Argon for cleaning the surface (f) in the positive ion mode (MC max counts per pixel, TC total counts).

### *3.2 Electrochemical properties of the NBD films*

The electrochemistry of the as-grown NBD and BDD films was then investigated and compared. Both inner and outer redox systems were used, namely 1 mM $[Ru(NH_3)_6]Cl_3$ (**Figure 3a**) and 1 mM $K_3[Fe(CN)_6]$ (**Figure 3b**) dissolved in 1 M KCl aqueous solution. For the $[Ru(NH_3)_6]Cl_3$ redox system (Figure 3a), the NBD electrode shows a higher peak current (e.g., a cathodic peak current, $I_c$ = 140.22 μA cm$^{-2}$) and a bigger difference of peak separation ($\Delta E_p$ = 72 mV) than the BDD electrode ($I_c$ = 131.08 μA cm$^{-2}$ and $\Delta E_p$ = 58 mV). For the $K_3[Fe(CN)_6]$ redox system (**Figure 3b**), $I_c$ rises from 145.5 μA cm$^{-2}$ on a BDD electrode to 162.34 μA cm$^{-2}$ on a NBD electrode. However, a NBD electrode shows a bigger $\Delta E_p$ (84 mV) than a BDD electrode (75 mV). As inner-sphere redox probes, the electrode kinetics of $[Fe(CN)_6]^{3-/4-}$ is known to be tightly related to surface terminations or surface functional groups of a diamond electrode [42]. Different from the $[Fe(CN)_6]^{3-/4-}$ inner-sphere redox system, the electron transfer and the electrode kinetic of the outer-sphere $[Ru(NH_3)_6]^{3+/2+}$ redox system is influenced mainly by the carrier density (e.g., the amount of sp$^2$ carbon species) of the diamond films [17, 45]. According to the growth parameters, the NBD and BDD films feature high boron densities or low electricity that is favorable for fast electron transfer processes [16]. However, a higher amount of boron atoms is expected to be doped in the NBD film than that in a BDD film. This originates from the "enhanced incorporation" effect of nitrogen in the gas mixture. The XPS results showed that the NBD film is enriched in C=O bonds compared to the BDD film. These surface oxygen groups on the electrodes thus block electrochemical active sites of the NBD electrode and/or bring more repulsive force for the negatively charged $[Fe(CN)_6]^{3-/4-}$ redox probes to interact with the NBD electrode. The electron transfer process of $[Fe(CN)_6]^{3-/4-}$ redox probes is thus inhibited on the NBD surface, eventually leading to reduced peak currents. On the other hand, the increased amount of $sp^2$ species after the nitrogen incorporation into a BDD film leads to the decrease of carrier density that promotes the electron transfer of outer-sphere $[Ru(NH_3)_6]^{3+/2+}$ redox probes and finally more pronounced peak currents.

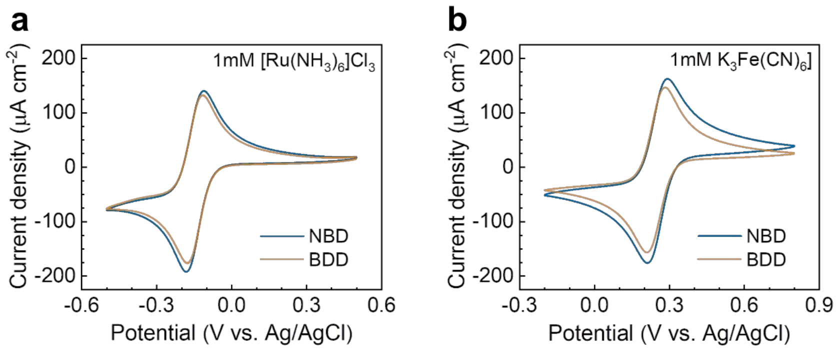


**Figure 3**. Cyclic voltammograms of the NBD and BDD electrodes at a scan rate of 0.1 V $s^{-1}$ in (a) 1 mM $[Ru(NH_3)_3]Cl_3$ and (b) 1 mM $K_3Fe(CN)_3$ dissolved in 0.5 M KCl solution.

*3.3 Electrochemical applications of the NBD films*

To explore the electrochemical applications of such NBD films, they were utilized as the capacitor electrodes for the supercapacitor assembly. In such a case study, these NBD films were wet-chemically treated since such post-treatment improved much their wettability. Here, a redox-electrolyte enhanced supercapacitor was fabricated [4, 46]. From the CVs of the NBD and BDD films recorded at different scan rates (**Figure S6**), one can notice stable $\Delta E_p$ and $I_c$ at all scan rates. These results indicate the perfect reversibility of the NBD and BDD films or these diamond capacitor electrodes in such an electrolyte. Notice here that the $\Delta E_p$ values in **Figure S5** are different from those in **Figure 3b**, although the used redox electrolytes are same. This is because these diamond electrodes in **Figure S5** were wet-chemically treated, while those in **Figure 3b** were the as-grown diamond films. In other words, different surface terminations on these electrodes affect significantly the kinetics of redox reactions on these diamond electrodes [15-17]. The estimated capacitances of the NBD electrode are 87.7, 66.8, 39.4, and 26.8 mF cm$^{-2}$ at the scan rates of 10, 20, 50, and 100 mV s$^{-1}$, respectively. Meanwhile, the galvanostatic charge/discharge (GCD) curves of the NBD and BDD electrodes (**Figure 4a**) also reveal good reversibility, as confirmed from the almost equal charge and discharge times in these GCD curves. The calculated capacitances of the NBD electrode (**Figure 4b**) are 98.9, 56, 28.7, and 14.9 mF s$^{-1}$ at the current densities of 1, 2, 4, and 8 mA cm$^{-2}$, respectively. They are higher than those of a BDD electrodes: 71.3, 39.7, 20.8, and 10.7 mF s$^{-1}$ at the current densities of 1, 2, 4, and 8 mA cm$^{-2}$, respectively. The capacitive contribution of the NBD and BDD electrodes were further calculated to explore the difference of the reaction kinetics between two capacitor electrodes. **Figure 4c** presents the contribution ratios of capacitive-controlled and diffusion-controlled processes on the NBD and BDD electrodes. Both exhibit an increased ratio of capacitive contribution with the enlargement of scan rate. Specifically, the NBD electrode shows a higher capacitive contribution ratio than a BDD electrode. This reveals the underlying essence of the better rate performance of the NBD electrode. The enhanced capacitance of the NBD electrode is because the incorporation of nitrogen and boron atoms into diamond modulates the surface polarities and the electronic of materials [47, 48]. For example, the charge-transfer resistance of the NBD film, as estimated from its Nyquist plots (**Figure S7**) is 94 Ω, which is smaller than that (143 Ω) of the BDD film.

The long-term cycling stability of the BDD and NBD electrodes was further tested at the current density of as high as 8 mA cm$^{-2}$. Although the NBD electrode exhibits a higher capacitance than a BDD electrode, both electrodes show the similar cycling stability even after 10000 GCD cycles (**Figure 4d** and **Figure S8**). All these results confirm the suitability of employing the NBD film for electrochemical energy storage applications. Note that the surface of the post-treated NBD electrode is possible to be re-activated electrochemically or by use of a plasma technique. The studies on the effect of the surface terminations of the NBD electrode on their capacitive performance are currently undergoing in our lab.

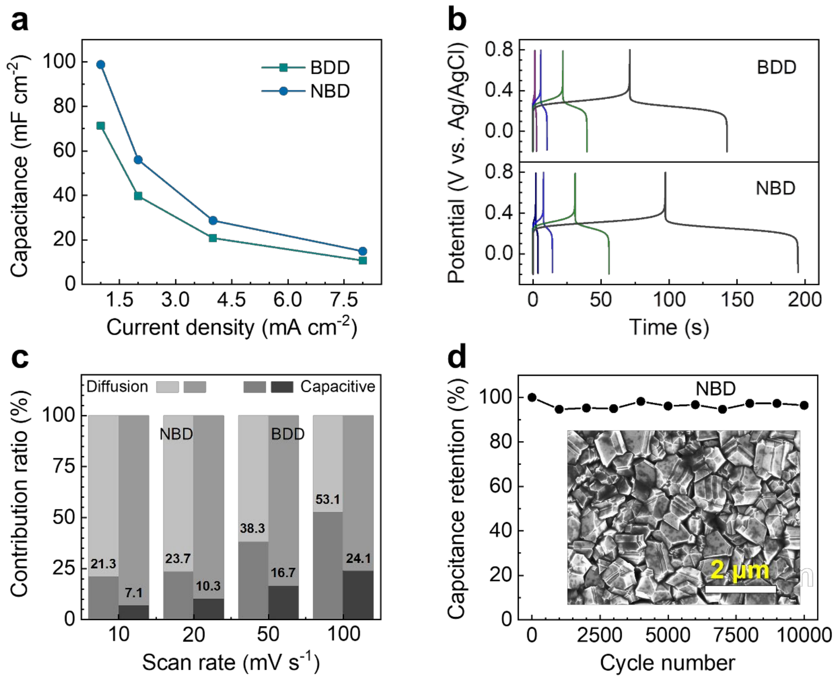


**Figure 4**. Capacitive performance of the NBD and BDD electrodes in 0.05 M $Fe(CN)_6^{3-/4-}$+ 1.0 M $NaSO_4$: (a) the GCD curves at the current densities of 1, 2, 4, and 8 mA $cm^{-2}$; (b) the variation of the specific capacitances with the current densities; (c) the contribution ratios of capacitive and diffusion capacity as a function of scan rates; (d) the capacitance retention at a current density of 8 mA $cm^{-2}$. The inset shows the SEM image of the post-treated NBD electrode after 10000 GCD cycles.

## 4. Conclusion

The electrochemistry of nitrogen and boron bi-element incorporated diamond film is explored. The NBD electrode reveals better electrochemical responses in both inner-sphere $[Fe(CN)_6]^{3-/4-}$ and outer-sphere $[Ru(NH_3)_6]^{3+/2+}$ redox systems, when compared to a BDD electrode. The improved electrochemical performance of the NBD film is related to the enrichment of C=O bonds and the increase amount of $sp^2$ species on the NBD film. The bigger capacitance of the NBD electrode than that of a BDD mainly stems from that the incorporation of nitrogen and boron atoms into the diamond modulates the surface polarities and the electronic structures of diamond. Moreover, a higher amount of boron atoms in the NBD film is expected than that in the BDD film, due to the "enhanced incorporation" effect of nitrogen in the gas mixture. Such an enhanced capacitance of the NBD electrode extends its potential applications for electrochemical energy storage. Future work has to be conducted on the effect of the densities of incorporated atoms in the NBD film and the surface terminations of the NBD film on the electrochemistry of these bi-element incorporated diamond films. Their further electrochemical energy storage (e.g., for SCs and batteries) and catalytic applications (e.g., for water splitting and $CO_2$ reduction, nitrogen fixation) can be tried. In summary, this work provides a new electrode material for future electrochemical applications.

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# Supporting Information

## Electrochemistry of Nitrogen and Boron Bi-element Incorporated Diamond Films

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## Supporting Tables

**Table S1**. CVD growth parameters of the boron-doped diamond (BDD) films as well as the nitrogen and boron bi-element incorporated diamond (NBD) films.

| | BDD | NBD |
|---|---|---|
| **Incubation** | | |
| Forward power (kW) | 4.8 | 4.8 |
| Chamber pressure (Torr) | 45 | 45 |
| Duration times (min) | 7 | 7 |
| $CH_4$ (sccm) | 15 | 15 |
| $H_2$ (sccm) | 185 | 82 |
| TMB (sccm~~ppm~~)* | 0.2~~2000~~ | 0.4~~2000~~ |
| $N_2$ (sccm) | | 3 |
| **Growth** | | |
| Forward power (kW) | 4.8 | 4.8 |
| Chamber pressure (Torr) | 45 | 45 |
| Duration times (min) | 1435 | 1203 |
| $CH_4$ (sccm) | 3 | 3 |
| $H_2$ (sccm) | 277 | 254 |
| TMB (sccm) | 20 | 40 |
| $N_2$ (sccm) | | 3 |

*** The TMB flow has been calculated based on total flow of gas mix containing 2000ppm TMB diluted in H2.**

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**Table S2**. Relative abundance of the carbon components in the BDD and NBD films.*

| | $sp^2$ C | $sp^3$ C | C-O | C=O | C-N | $sp^2$ C/$sp^3$ C |
|---|---|---|---|---|---|---|
| NBD | 3.7 | 31.0 | 34.0 | 28.9 | 2.4 | 11.9 |
| BDD | 2.0 | 52.0 | 32.3 | 13.6 | | 3.8 |

* These atomic ratios were estimated from their high resolution C1s XPS spectra

## Supporting Figures

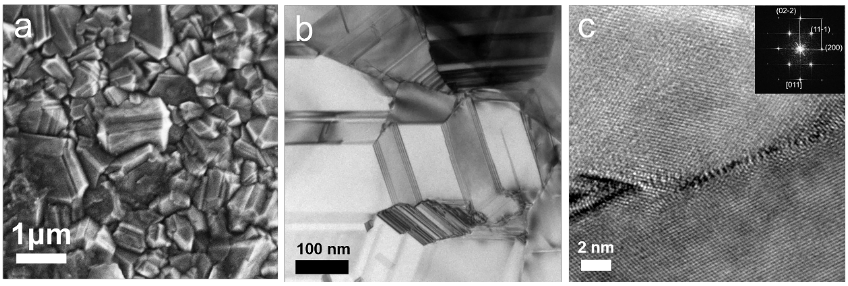


**Figure S1**. (a) SEM, (b) low-magnification TEM and (c) HRTEM images of the BDD film. The inset in (c) is the corresponding fast Fourier transformation (FFT) of the HRTEM image.

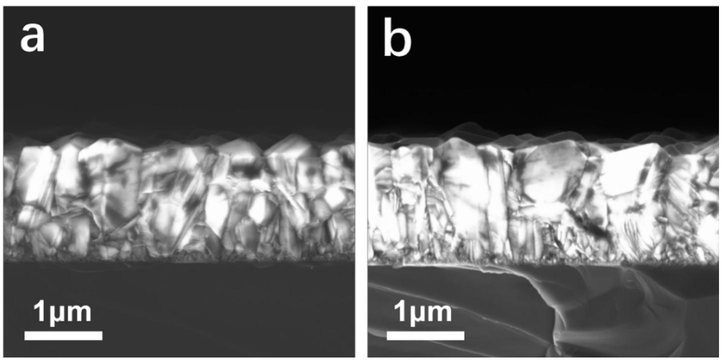


**Figure S2**. The cross-sectional SEM images of the (a) NBD and (b) BDD films.

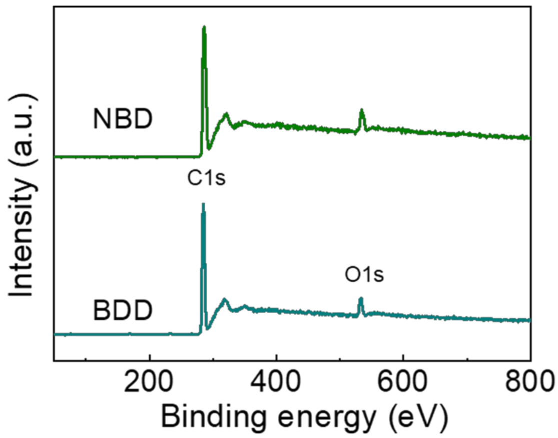


**Figure S3.** XPS survey spectra for the NBD and BDD films.

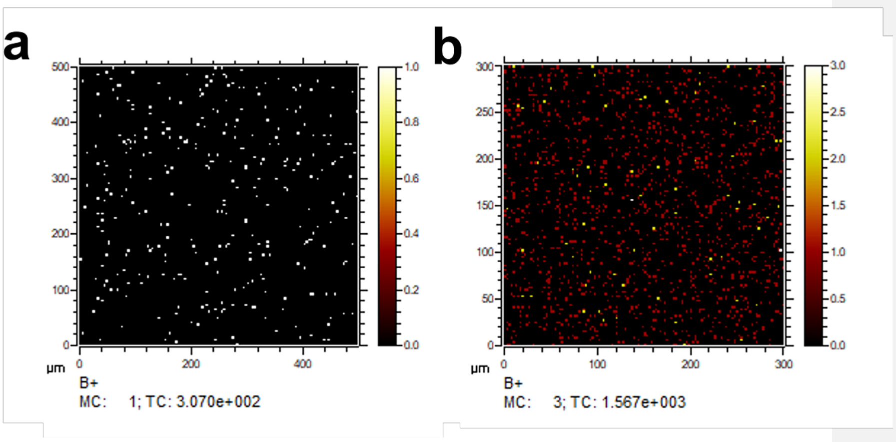


**Figure S4**. SIMS mapping of boron atoms doped in the BDD film (a) direct and (b) after 30 sec sputtering with Argon for cleaning the surface in the positive mode (MC max counts per pixel, TC total counts).

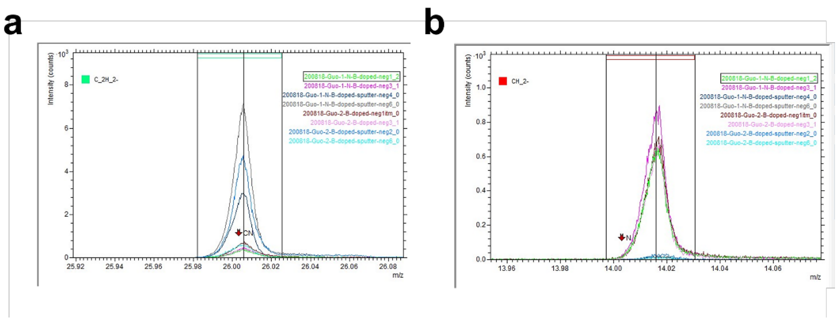


**Figure S5.** Spectra of the (a) NBD and (b) BDD films in the negative mode. The *m/z* ratios of N, $CH_2$, CN, and $C_2H_2$ are 14,003, 14,0162, 26,0036 and 26,0162, respectively.

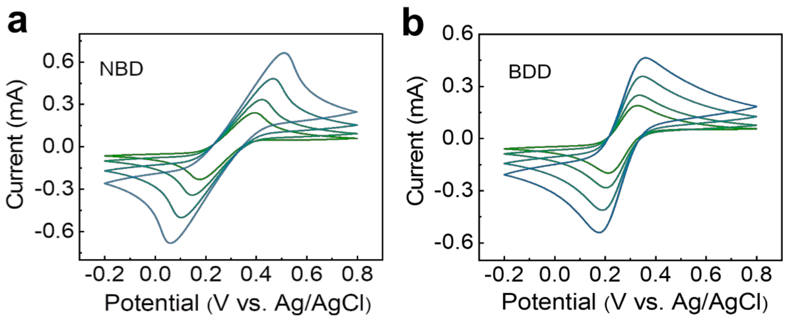


**Figure S6**. CVs of 0.05 M $Fe(CN)_6^{3-/4-}$ in 1.0 M $NaSO_4$ on the (a) NBD and (b) BDD electrodes at the scan rates of 100, 50, 20, and 10 mV $s^{-1}$.

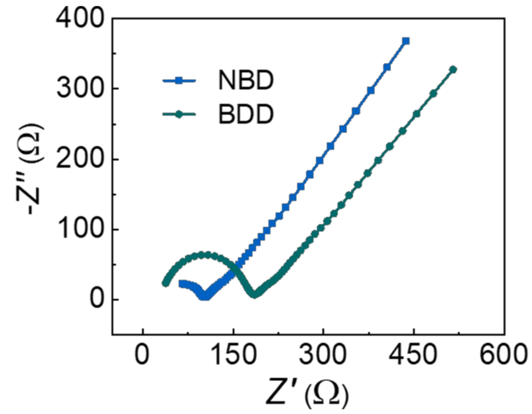


**Figure S7**. Nyquist plots of the NBD and BDD electrodes in 1.0 M $NaSO_4$ solution containing 0.05 M $Fe(CN)_6^{3-/4-}$.

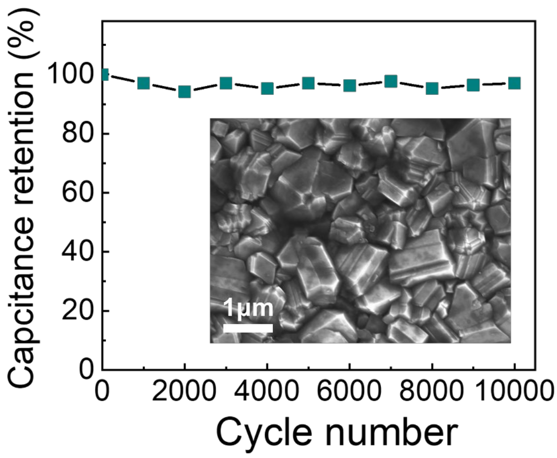


**Figure S8**. Capacitance retention of a BDD electrode at a current density of 8 mA $cm^{-2}$. The inset shows the SEM image of a BDD electrode after 10000 GCD cycles.